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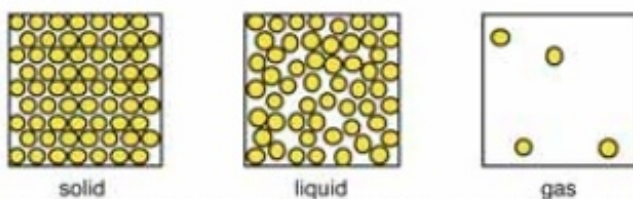
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Answers

Chapter 1

- a) melting
b) freezing
c) subliming / sublimation
d) subliming / sublimation

2. a)



Note: Solids should have regularly packed particles touching. Liquids should have most of the particles touching at least some of their neighbours, but with gaps here and there, and no regularity. Gases should have the particles well spaced.

b) Solids: vibration around a fixed point. Liquids: particles can move around into vacant spaces, but with some difficulty because of the relatively close packing.

c) Evaporation: Some faster moving particles break away from the surface of the liquid. Boiling: Attractive forces are broken throughout the liquid to produce bubbles of vapour.

d) In a sealed container, vapour particles in the space above the liquid return and stick to the surface of the liquid at the same rate as liquid particles are evaporating.

3. a) A - gas ; B - liquid; C - solid; D - liquid; E - solid

b) C. It sublimes at a very high temperature and so takes the most heat energy to break the attractions between the particles to form a gas.

c) A. It turns to a gas at the lowest temperature and so takes the least heat energy to break the attractions between the particles to form a gas.

d) A, because it is a gas.

e) It sublimes. That means that it turns straight from a solid to a gas and so there is never any liquid to boil.

f) D. It has the lower boiling point of the two liquids (B and D), and therefore the weaker attractions between its particles. A higher proportion of its particles will have enough energy to escape from the surface.

4. a) The ammonia and hydrogen chloride particles have to diffuse through the air in the tube, colliding with air particles all the way.

b) (i) Its particles will move faster.

(ii) It would take slightly longer for the white ring to form, because the gas particles would be moving more slowly at the lower temperature.

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